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(21) International Application Number: PCT/IT99/00225 (22) International Filing Date: 19 July 1999 (19.07.99) (30) Priority Data: RE99A000028 5 March 1999 (05.03.99) IT (71) Applicant (for all designated States except US): AGRISTUDIO S.R.L. [IT/IT]; Via Gramsci 104/F, I-42100 Reggio Emilia (IT). (72) Inventor; and (75) Inventor/Applicant (for US only): CIRIBOLLA, Antonio [IT/IT]; Via Martiri della Bettola, 96, I-42100 Reggio Emilia (IT). (74) Agent: COLLI, Alfonso; Bugnion S.P.A., Via Roma, 38, I-42100 Reggio Emilia (IT).		(81) Designated States: AE, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ZA, ZW, ARIPO patent (GH, GM, KE, LS, MW, SD, SL, SZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG). Published <i>With international search report.</i>
(54) Title: CHELATED FEED ADDITIVE AND METHOD OF PREPARATION (57) Abstract Feed additive for agro-zootechnical use, in particular for alimentary use in the zootechnical sector, consisting of a chelate obtained by the reaction of methionine hydroxy analogue with the carbonate of a bivalent metal. The reaction is free from undesirable by-products, and the product is stable and effective in improving the main growth factors of the animal.		

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- 1 -

Description

CHELATED FEED ADDITIVE AND METHOD OF PREPARATION

Technical Field

The present invention relates to a chelated additive for agro-zootechnical use in the broad sense, and in particular for alimentary use in the zootechnical sector, and the associated method for obtaining said additive.

- 5 A chelate is a compound which is obtained from an organic molecule (in the case in question an amino acid or a peptide chain) and a metal ion by means of strong coordination bonds.

Background art

- 10 These compounds are widely used both in the agronomic field and in the zootechnical field since their usefulness is derived from the biological action of the metal involved, which acts as activator in a great number of enzymatic reactions and as regulator in various
- 15 metabolic functions in all living organisms.

The presence of the bond with an amino acid material favours the absorption, availability and use of the metal since it is transported by the organic component to all

- 2 -

areas of the organism.

Within the zootechnical sector in particular, these compounds are used in various breeding or rearing sectors to increase and strengthen the normal metabolic and functional activities of organs and apparatuses, with
5 considerable positive effects on the capacities for productive growth of the animals.

The chelates currently present on the international market may be classified in various categories.

10 The best known category is that which groups together the metal ions which form a complex with specific single amino acids. They are well definable compounds in terms of chemical composition and are easily assimilated and available for the organism.

15 A second category is composed of micronutrients which form a complex with single amino acids which, even though not well definable in this case, possess characteristics similar to those of the preceding group.

During production of these compounds, it was noted,
20 moreover, that some soluble metal salts are difficult to chelate or leave undesirable residual products such as, for example, phosphoric acid or hydrochloric acid.

A third category contains metal ions which form a complex

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During production of these compounds, it was noted,
20 moreover, that some soluble metal salts are difficult to chelate or leave undesirable residual products such as, for example, phosphoric acid or hydrochloric acid.

A third category contains metal ions which form a complex

- 3 -

with several amino acids, preferably a maximum of three up to a molecular weight of 800 daltons.

The greater size of the molecule, however, results in a weaker bonding strength and greater difficulty of absorption.

A further category is composed of proteins which are obtained from the combination of a metal ion with a peptide chain. These compounds, the amino acid composition of which is not known, may interact with other substances contained in the food and therefore be of limited availability for the animal.

Finally, there are metals which form a complex with polysaccharide chains, in which the organic component has little effect on the real nutritional requirements of the animal and the metals are combined with organic acids, which are very soluble, but may easily dissociate in the intestines, with adverse effects in terms of lowering of the intestinal pH.

Disclosure of the invention

The object of the present invention is to obtain an additive for agro-zootechnical use and in particular for alimentary use in the zootechnical sector, in which a metal, which may also be classified among the

- 4 -

micronutrients, is bonded to an amino acid via a strong chelating bond, this product being not only effective, but also stable and easily reproducible, and free from undesirable reaction by-products.

- 5 The objects are achieved by using a methionine in solution with an acid function (methionine hydroxy analogue) and reacting it with a carbonate of a bivalent metal, starting the reaction of the two components in particular with less methionine, relative to the
10 carbonate.

Experiments carried out using the product thus obtained on dairy cattle have shown that the invention has positive and important effects on the animals.

- Two groups of animals were prepared, one test group
15 supplied with ordinary metal salts, and one experimental group supplied with the same quantity of chelated metal ions, but with the methionine hydroxy analogue according to the invention, in doses of about 10 g per head of cattle per day.

- 20 At the end of the test, the following results were observed in the treated group: a significant improvement in all the reproductive factors (fertility, succesful pregnancies, more rapid return to heat and improved ovary

- 5 -

activity) and reduction in infections of the uterus and cases of mastitis following an improvement in the immune state of the animal, which allowed the latter to withstand effectively all the problems of an infective and metabolic nature which are typical of the postnatal period. Similar experiments were carried out on some meat cattle stock, distinguishing a test group from an analysis group and supplying the said groups with identical quantities of metal ions, one in salt form and the other in chelated form for about three months, in doses of about 10 g per head of cattle per day.

In this latter test, too, surprisingly a greater resistance to environmental stresses on the part of the animals, in the form of improved productive performances (growth and conversion factors, body growth), was noted. A reduction in podalic pathologies (laminitis, ulcers and interdigital dermatitis) was also detected.

The invention will emerge more clearly from the description of a working example illustrated below.

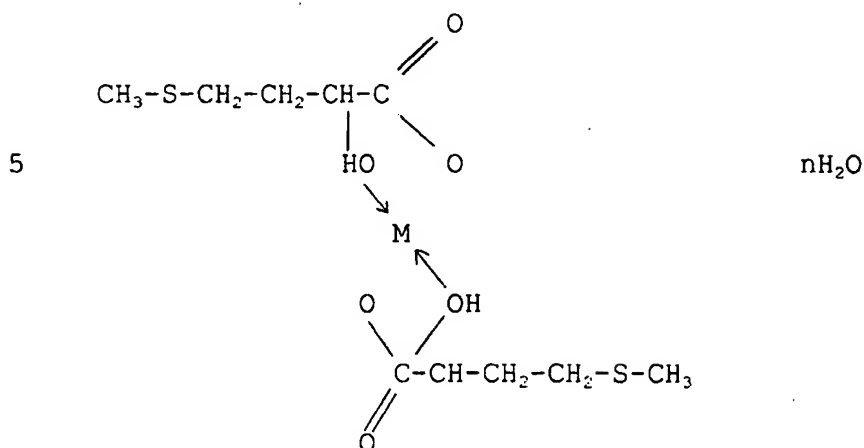
Detailed description of the preferred embodiments

Example No. 1

The structural formula of a chelate according to the invention, obtained from the reaction of a methionine

- 6 -

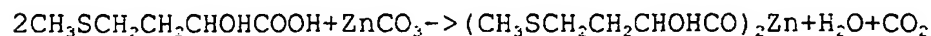
hydroxy analogue with the carbonate of a bivalent metal,
in the presence of water, is shown:



10 and in short $(\text{CH}_3\text{SCH}_2\text{CH}_2\text{CHOHCO})_2\text{M}\cdot n\text{H}_2\text{O}$

where M may be a bivalent metal and is in general Zn^{2+} ,
 Cu^{2+} , Co^{2+} , Mn^{2+} , Ca^{2+} or Mg^{2+} .

Said compound was prepared using in particular zinc
carbonate; it is obtained from the reaction of two
15 molecules of methionine hydroxy analogue + one of ZnCO_3 :



where, since the reaction takes place in an aqueous
solution, water of crystallization may be present in the
dried chelate.

20 The reaction takes place in the presence of a small
amount of demineralized water, at temperatures of between
25°C and 35°C and at ambient pressure, with the evolution
of CO_2 .

- 7 -

Although the reaction is slightly exothermic, slight heating up to 40°C may be appropriate in order to accelerate the process and facilitate the evaporation of the excess water.

- 5 In the case described, the dried chelate is in the form of a white powder, with a melting point of about 248°C and a solubility in water of about 20g/l at ambient temperature.

The infrared vibrational absorption spectrum indicates
10 that the chelating process has taken place: in fact, the spectrum exhibits a series of characteristic bands, in accordance with the structure described above. The strongest bands are indicated below (values in cm^{-1}): 3219 (OH stretching); 2918 (CH stretching); 1591 (COO
15 asymmetrical stretching); 1422 (COO symmetrical stretching + OH bending); 1307 (CH bending); 1088 (C-OH stretching).

The method of preparation of the product is as follows:
One mole of zinc carbonate is poured, together with
20 demineralized water slightly heated to about 35°C, in a sufficient quantity to dissolve the zinc carbonate, into a mixer or a reactor provided with a mixer.

The optimum temperature range is in any case between 20°C

- 8 -

and 40°C.

The mixer is turned on and the methionine hydroxy analogue is added slowly so as to avoid the undesirable effects of a strong evolution of CO₂ and overheating of the mixture. The stoichiometric ratio is two moles of methionine hydroxy analogue to one mole of zinc carbonate; said ratio being reached slowly by means of successive additions of methionine while mixing is continued for a duration of between one and two hours so as to favour the generation of CO₂ and evaporation of the excess water.

During the entire reaction, the methionine/metal carbonate stoichiometric ratio is less than the theoretical ratio; approaching the theoretical value asymptotically at the end of the reaction.

Once the product has reached a paste-like consistency, drying may be performed by means of the normal drying processes which are known today, or by adding adsorbent products (silica, corn-cob flour).

In this way a stable product, which is pure in that it is devoid of reaction by-products, is obtained.

Moreover this modified amino acid offers considerable advantages over other common amino acids in that:

- 9 -

- * it can be more easily assimilated in the duodenum and is less susceptible to degradation of the ruminal bacterial flora;
 - * it may be rapidly detected in the blood and is therefore available for production;
 - * it can be mixed uniformly without the risk of compaction;
 - * it is compatible with vitamins, fats and minerals present in the diet;
 - * it increases the efficiency of anti-mould products owing to its acid characteristics;
 - * absorption thereof in the intestines does not require energy which is dissipated in the form of a dynamic specific action.
- 15 The chelate obtained was prepared using simple apparatuses already present in any normal additive-manufacturing industry.
- Finally the costs for these operations are very low.
- Possible variants:
- 20 The possible variants are obtained by simply exchanging the bivalent metal (cuprous ion, zinc, manganous ion, cobalt, ferrous ion) chelated in each case with the organic molecule.

- 10 -

For example, if, for the sake of brevity, HMetOH designates the methionine hydroxy analogue where H is the exchangeable acid, the solid products which can be obtained by the process illustrated above may be indicated by $M(\text{MetOH})_2 \cdot n\text{H}_2\text{O}$ where M is the bivalent metal cation of a carbonate.

When M is Co^{2+} , a pink powder with a melting point $> 310^\circ\text{C}$ is obtained.

When M is Mn^{2+} , a light brown powder with a melting point $> 310^\circ\text{C}$ is obtained.

When M is Cu^{2+} , a light green powder with a melting point of 202°C is obtained.

Each metal ion has particular characteristics with regard to given enzymatic and metabolic systems of the animal or the plant to which it is administered, whereby the advantages of administering it in the form of a chelate bonded to a methionine hydroxy analogue remain unaffected.

Illustrated with regard to applications in the zootechnical sector, the additive according to the invention may also find widespread application in the agronomic sector, where the presence of the bond between the bivalent metal and an amino acid material favours the

- 11 -

absorption, availability and use of the metal since it is transported by the organic component to the vegetable tissue.

- 12 -

Claims

1. Additive for agro-zootechnical use and in particular for alimentary use in the zootechnical sector, consisting of a chelate produced by the combination of an organic molecule with a metal, characterized in that the organic part consists of a methionine in solution with an acid function (methionine hydroxy analogue) and the inorganic part consists of a bivalent metal in the form of a carbonate salt.
2. Additive for zootechnical use according to Claim 1, characterized in that it is represented by the formula $(\text{CH}_3\text{SCH}_2\text{CH}_2\text{CHOHCO})_2\text{M} \cdot n\text{H}_2\text{O}$ where M may be a bivalent metal and in particular Zn^{2+} , Cu^{2+} , Co^{2+} , Mn^{2+} , Ca^{2+} or Mg^{2+} .
3. Additive for zootechnical use according to Claim 1, characterized in that it consists of an organic Zn salt having the following approximate characteristics: white powder, with a melting point of about 248°C , solubility in water of about 20g/l at ambient temperature; and infrared vibrational absorption spectrum having a series of characteristic bands (values in cm^{-1}): 3219 (OH stretching); 2918 (CH stretching); 1591 (COO asymmetrical

- 13 -

stretching); 1422 (COO symmetrical stretching + OH bending); 1307 (CH bending); 1088 (C-OH stretching).

4. Method for obtaining an additive for agro-
5 zootechnical use, characterized in that it consists of the following processes:

- a) dissolution of Zn carbonate in water;
- b) mixing;
- c) slow and continuous addition of methionine hydroxy
10 analogue, until the stoichiometric ratio of two moles of methionine to one mole of Zn carbonate is reached;
- d) final drying.

5. Method according to the preceding Claim 4,
15 characterized in that it is conducted in the temperature range from 20°C to 40°C.

6. Method for obtaining an additive for zootechnical use according to Claim 1, characterized in that, during
20 the entire reaction, the methionine/metal carbonate stoichiometric ratio is less than the theoretical ratio, approaching the theoretical value asymptotically at the end of the reaction.

INTERNATIONAL SEARCH REPORT

International Application No

PCT/IT 99/00225

A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 A23K1/16 C07C323/52 C07C319/20 A23K1/175

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 A23K C07C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 4 579 962 A (TAKANO MASA HARU) 1 April 1986 (1986-04-01) column 1, line 25-29 column 2, line 50-53 column 4, line 5-7 example 8 claims 1,6,8	1-3
A	see above citations column 3, line 1-4	4-6
P,X	PATENT ABSTRACTS OF JAPAN vol. 1999, no. 08, 30 June 1999 (1999-06-30) & JP 11 075885 A (NIPPON SODA CO LTD), 23 March 1999 (1999-03-23) abstract	1,2
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☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

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Date of the actual completion of the international search

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19/11/1999

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INTERNATIONAL SEARCH REPORT

International Application No
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A	EP 0 049 057 A (MONSANTO CO) 7 April 1982 (1982-04-07) page 3, line 19-21 example 1 claim 6	4-6
A	WO 96 36598 A (ZINPRO CORP) 21 November 1996 (1996-11-21) page 9; example 5	4-6

INTERNATIONAL SEARCH REPORT

Information on patent family members

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